

Name: \_\_\_\_\_

Period: \_\_\_\_\_

### The Periodic Table Exam Review

Who designed the first Periodic Table and how was it arranged? \_\_\_\_\_

Who designed the modern Periodic Table and how is it arranged? \_\_\_\_\_

Write the name that is given to the following groups of elements:

Group A Elements: \_\_\_\_\_ Group B Elements: \_\_\_\_\_

Group 1A Elements: \_\_\_\_\_ Group 2A Elements: \_\_\_\_\_

Group 7A Elements: \_\_\_\_\_ Group 8A Elements: \_\_\_\_\_

The two rows of elements at the bottom of the Periodic Table: \_\_\_\_\_

Fill in the blanks with s, p, d, or f block:

Inner transition metals: \_\_\_\_\_

Groups 1A&2A: \_\_\_\_\_

Transition metals: \_\_\_\_\_

Groups 3A-8A(or 0): \_\_\_\_\_

Arrange the elements Li, Cs, and F in order of increasing atomic sizes: \_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_

In each pair, circle which radius is larger: K or  $K^+$  Br or  $Br^-$

Define ionization energy (IE): \_\_\_\_\_

(increases or decreases) In general, ionization energy \_\_\_\_\_ as you go across a period and \_\_\_\_\_ as you go down a group.

Circle the element that has the lower IE: Cl or Na

Define electronegativity: \_\_\_\_\_

Rank the elements B, F, Cs, and K from lowest electronegativity to highest: \_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_

In general, nonmetals have (low or high) electronegativity and tend to (lose or gain) electrons when they form ions.

Rank the elements Al, Na, and O in order of increasing ionization energy: \_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_

Circle which element or ion would be smaller:

Ca or  $Ca^{2+}$

O or  $O^{2-}$

Br or  $Br^-$

Na or  $Na^+$

As you move from left to right across the periodic table, what happens to the size of an atom?

\_\_\_\_\_

