

REPORT
MOLECULAR MODELS EXP.

NAME _____

SECTION _____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
SCl_2					
PCl_3					
O_3					
SiH_2Cl_2					

REPORT FOR MOLECULAR MODELS EXP. (cont.) NAME _____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
PCl_5					
ICl_2^-					
SF_6					
XeF_5^-					

REPORT FOR MOLECULAR MODELS EXP. (cont.) NAME _____

Compound	Lewis Structure	Central Atom Hybrid	Electron Pair Geometry	Molecular Geometry	Polarity
CO					
PCl ₃					
SF ₄					
NO ₃ ⁻					

QUESTIONS FOR MOLECULAR MODELS EXP. NAME _____

1. Answer each of the following for the Nitrate ion (NO_3^-).

a. Provide the Lewis structures (including resonance forms) for the Nitrate ion. In one of the structures label the hybridization of the Nitrogen and each of the Oxygens.

b. What are the electron pair and molecular geometries?

electron pair _____

molecular _____

c. Is this a polar or nonpolar ion?

2. The Nitrite ion is slightly different than the Nitrate ion. What are the electron pair and molecular geometries for Nitrite ion? What is its polarity? Include the Lewis structure for this ion in your response.