Name	
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Date _____ Pd ____

Chemistry – Unit 5 & 6 Review Extra Credit Empirical Formula & Molar Mass Worksheet Directions: Turn the following into balanced equations by filling in the blanks with the correct

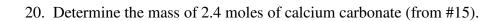
coefficients, formulas of ions or solids, and names.

Cation Anion	Formula	Name
1 Na ⁺ + Br ⁻	→	
2 Cu ⁺ + SO ₄ ²⁻	→	
3Pb ²⁺ +Cl ⁻	→	
4. $\underline{\hspace{1cm}} K^+ + \underline{\hspace{1cm}} S^{2-}$	→	
5 Sn ²⁺ + F	→	
6 +	$\rightarrow BaI_2$	
7 +	\rightarrow AlCl ₃	
8 +	\rightarrow Mg(NO ₃) ₂	
9 +	$\rightarrow KC_2H_3O_2$	
10 +	$\rightarrow (NH_4)_2SO_3$	
11 +	→	silver oxide
12 +	→	iron(III) sulfide
13 +	→	copper(II) nitrate
14 +	→	magnesium chloride
15 +	→	calcium carbonate
16. $Mg^{2+} + NO_2^{-}$	→	
17 Cu ²⁺ + OH ⁻	→	

18. \longrightarrow K₂CrO₄

Answer the following questions. Show all work, labeled, with units.

19.	Calculate the molar mass of iron(III) sulfide (the compound in question $\#12$)



21. How many grams of BaI₂ are in 3.9 moles? (from #6)

22. How many moles of AlCl₃ formula units are contained in a 20.0 g sample? (from #7)