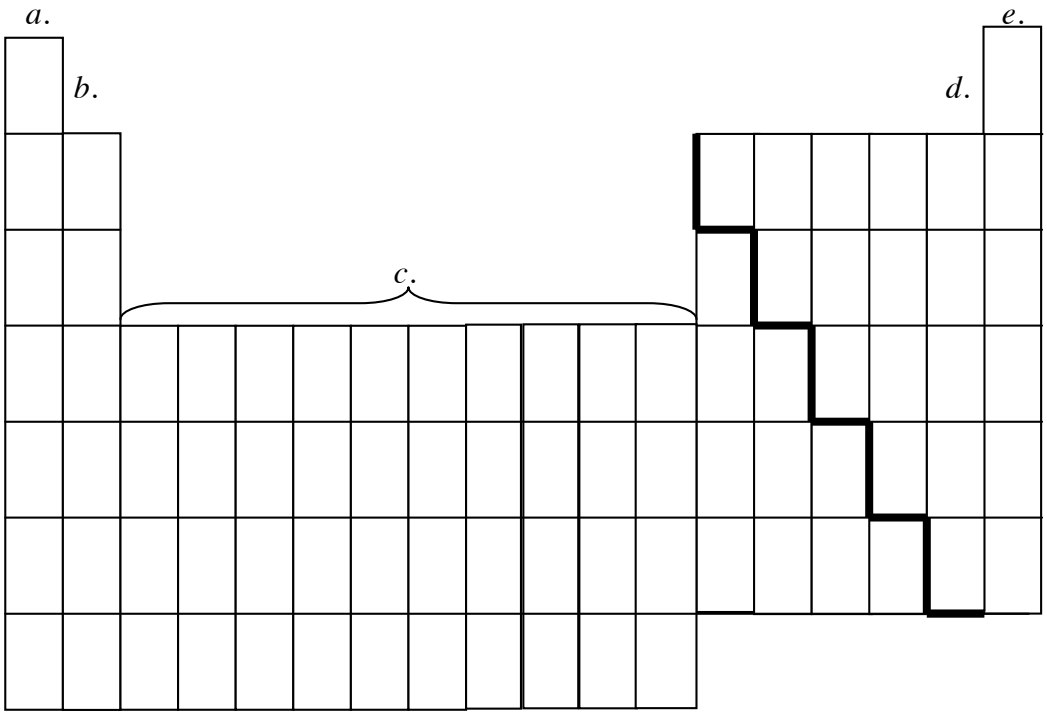
Name: _____

Date: _____ Per: _____

Chemistry
Periodicity WS

I. *Identify the following groups of the periodic table and give their number of valence electrons:*



Group	Group Name	# of valence electrons	ion formed
a			
b			
c			
d			
e			

II. *Define the following:*

1. periodicity _____
2. electronegativity _____
3. ion _____
4. ionization energy _____

III. *Answer the following questions about periodic trends:*

5. What tends to happen to atomic radius as you go from left to right on the periodic table?

_____

6. What tends to happen to electronegativity as you go from top to bottom on the periodic table?

_____

7. Will boron or aluminum have a larger radius? Why?

_____

8. Will neon or sodium have greater ionization energy? Why?

_____

IV. *Answer the following questions about the element aluminum:*

9. What are three characteristic properties of metals?

_____

10. Aluminum is a metal, but it doesn't exhibit all the normal properties of metals. What property is missing from aluminum? _____

11. Aluminum is next to a dark, jagged line on the periodic table. What is this line called and what does it mean? _____

_____

V. *Answer the following questions about the element magnesium:*

12. What group or family is magnesium found in? _____

13. What is another element from this group? _____

14. How many electrons does magnesium have in its highest energy level? _____

15. What other elements have the same number of outer electrons as magnesium?

_____

16. Knowing their ionization energies, is sodium or magnesium more likely to form an ion? Why?

_____