Bond Polarity Worksheet
Chemistry Chapter 8
2 points

Name: _____

Date: _____ Hour: _____

1. Use the electronegativity table on page 265 of your textbook to decide if the following bonds are ionic, polar covalent, or nonpolar covalent:

| | | Electronegativity Difference | Type of Bond | If polar, show partial charges |
|---|---|---|---|---|
| ex> | C - O | 0.89 | polar covalent | δ+ C–O δ- |
| a) | Cl - As | _____ | _____ | |
| b) | O - K | _____ | _____ | |
| c) | N - N | _____ | _____ | |
| d) | O - H | _____ | _____ | |
| e) | Si - S | _____ | _____ | |
| f) | N - P | _____ | _____ | |
| g) | Si - N | _____ | _____ | |
| h) | O - P | _____ | _____ | |
| i) | Mg - O | _____ | _____ | |

Use your notes or the information on pp. 265-267 in your textbook to answer the following:

2. What is the difference in the electrons in an ionic and covalent bond? _____

_____

_____

3. What is the difference in the electrons in a polar covalent and nonpolar covalent bond? ____

_____

_____

4. What does the lowercase Greek letter delta (δ) represent? _____

5. How do you pick the negative end or a polar covalent bond or ionic bond? _____

_____

_____