

Name: _____

ID: A

10. (EC) What is the empirical formula for a compound that is 53.3% O and 46.7% Si?
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|----------------------------|-------------------|
| a. Si_2O | c. SiO_2 |
| b. Si_2O_3 | d. SiO |
11. What is the percentage composition of CF_4 ?
- | | |
|---------------------|---------------------|
| a. 16.8% C, 83.2% F | c. 13.6% C, 86.4% F |
| b. 20% C, 80% F | d. 81% C, 19% F |

Numeric Response

12. The molar mass of aluminum is 26.98 g/mol and the molar mass of oxygen is 16.00 g/mol. Determine the molar mass of Al_2O_3 .
13. Calculate the number of moles in 10.05 grams of CaCO_3 (molar mass = 100.1g/mol).
14. What is the mass of 10.0 moles of carbon?
15. Calculate the molar mass of nitroglycerin, $\text{C}_3\text{H}_5(\text{NO}_3)_3$, the main ingredient of dynamite.
16. The percentage of hydrogen in NaOH is _____.