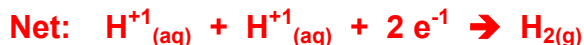


Write a balance equation and label the states of the reactants and products for each in each reaction. After writing the balanced equation, list the driving force and spectator ions, and write the net ionic equation(s) below the balanced equation. If there is no driving force for a reaction, write no reaction in place of the net ionic equation.

1. Silver acetate + sodium iodide  $\longrightarrow$



2. Nitric acid + potassium  $\longrightarrow$



3. Barium hydroxide + chloric acid  $\longrightarrow$



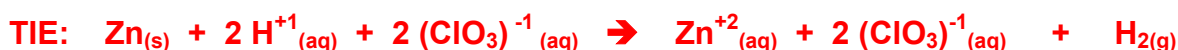
4. Ammonium carbonate + calcium nitrate  $\longrightarrow$



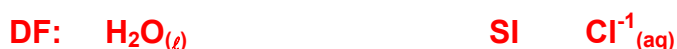
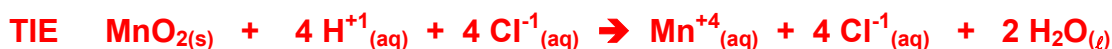
5. Acetic acid + potassium hydroxide  $\longrightarrow$



6. Zinc plus chloric acid  $\longrightarrow$



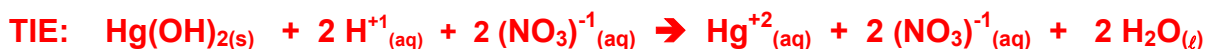
7. Manganese IV oxide + hydrochloric acid  $\longrightarrow$



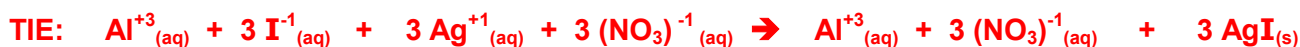
8. Carbonic acid + Iron (II) nitrate  $\longrightarrow$



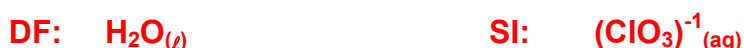
9. Mercury (II) hydroxide + nitric acid  $\longrightarrow$



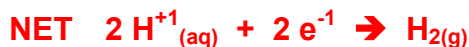
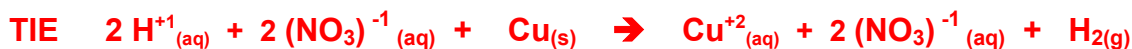
10. Aluminum iodide + silver nitrate  $\longrightarrow$



11. Chloric acid + magnesium oxide  $\longrightarrow$



12. Nitric acid + copper (II)  $\longrightarrow$



13. Hydrochloric acid + sodium hydroxide  $\longrightarrow$



14. Nitric acid + ammonia  $\longrightarrow$

