Experiment 7 The Chemistry of Oxygen: Basic and Acidic Oxides and the Periodic Table

Pre-Lab Quiz

Cours Date		Name Lab Section Number	
1.		how how the hydroxide and oxide compound of an element are related to each other by writing the reaction nat converts barium oxide to barium hydroxide.	
		$BaO(s) + \underline{\hspace{1cm}} \rightarrow \underline{\hspace{1cm}}$	
2.	a)	$Al_2O_3(s) + 3 H_2O(l) \rightarrow Al^{3+}(aq) + 6 OH^{-}(aq)$	
	b)	$Al_2O_3(s) + 7 H_2O(l) \rightarrow 2 Al(OH)_4^-(aq) + 2 H_3O^+(aq)$	
	Which	Which of the reactions, (a) or (b), shows the acidic nature of aluminum oxide?	
	Which	Which of the reactions, (a) or (b), shows the basic nature of aluminum oxide?	
	Toget	her, reactions (a) and (b) show the nature of aluminum oxide.	