

Name _____ Period _____ Date _____

Home Work 401: Intro to Periodic Table

- _____ In his periodic table, Mendeleev did not always list elements in order of increasing atomic mass because he wanted to group elements with similar
a. Properties
b. Atomic numbers
c. Densities
d. Colors
- _____ Mendeleev arranged the elements by increasing atomic:
a. Number
b. Mass
c. Color
d. Density
- _____ A few elements were out of place on Mendeleev's periodic table. Name the scientist who put them in the right place
a. Mendeleev
b. Democritus
c. Moseley
d. Dalton
- _____ Mosely discovered that elements with similar properties occurred at regular intervals when arranged in order of increasing...
a. Atomic mass
b. Density
c. Color
d. Atomic number
- _____ What is it about the elements in a column of the periodic table that makes them behave similarly?
a. Atomic number
b. Atomic mass
c. Electron distribution
d. Radioactivity
- _____ The discovery of the noble gases changed Mendeleev's periodic table by adding a:
a. Period
b. Series
c. Group
d. Sublevel
- _____ Then modern periodic law states that:
a. No two electrons can have the same spin in the same atomic orbital
b. The physical and chemical properties of the elements are functions of their atomic number
c. Electrons exhibit properties of both particles and waves
d. The chemical properties of elements can be grouped according to periodicity, by physical properties cannot
- _____ Which of the following is the noble gas distribution for arsenic?
a. $[Kr] 4s^2 3d^{10} 4p^3$
b. $[Ar] 4s^2 3d^{10} 4p^3$
c. $[Ge] 4p^3$
d. $[K] 4s^1 3d^{10} 4p^3$