

Name: _____

Date: _____

Hour: _____



Topic: Nomenclature

Content Standard(s):

Formula II Worksheet

Instructions- Answer the following questions on a separate sheet of paper. Make sure to show all appropriate work to justify your answer.

1. Use the periodic table to help you write the symbol for the ion that you would expect from each element below.

a. cesium b. barium c. gallium d. sulfur e. iodine
f. radium g. chlorine h. astatine i. oxygen j. argon

2. Write the chemical formula for each of the substances named.

a. lead (IV) chloride e. potassium oxide
b. iron (III) oxide f. gold (I) sulfide
c. nickel (II) nitride g. tin (IV) sulfide
d. manganese (IV) oxide h. uranium (IV) fluoride

3. Name the compounds below.

a. PbI_2 d. Co_3N_2
b. Li_3P e. Mg_3As_2
c. MgF_2 f. Pt_3P_4

4. Name the following molecular compounds.

a. CO_2 d. NO_2
b. CO e. Nl_3
c. CF_4 f. SiO_2

5. Write the formulas for the following:

a. carbon tetrachloride b. lead (IV) sulfide
c. iron (III) carbonate d. bismuth (V) phosphide
e. potassium bicarbonate f. barium phosphate

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Formula II Worksheet (Cont.)

Anion	Nitrate	Sulfite	Bicarbonate	Hydroxide	Dichromate
Cation	NO_3^-	SO_3^{2-}	HCO_3^-	OH^-	$\text{Cr}_2\text{O}_7^{2-}$
Lead Pb^{2+}	Example: Lead (II) nitrate $\text{Pb}(\text{NO}_3)_2$				
Sodium Na^+					
Silver Ag^+					
Ammonium NH_4^+					
Copper Cu^{2+}					