

Name: \_\_\_\_\_  
Hour: \_\_\_\_\_ Date: \_\_\_\_\_

## Chemistry: Oxidation Numbers and Ionic Compounds

Write the correct formula for the compound formed by each of the following pairs of ions.

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|---|-----------|
| 1. $\text{Na}^{1+}$ $\text{F}^{1-}$     | 1. _____  |
| 2. $\text{K}^{1+}$ $\text{S}^{2-}$      | 2. _____  |
| 3. $\text{Ni}^{2+}$ $\text{SO}_4^{2-}$  | 3. _____  |
| 4. $\text{Al}^{3+}$ $\text{O}^{2-}$     | 4. _____  |
| 5. $\text{Ca}^{2+}$ $\text{ClO}_3^{1-}$ | 5. _____  |
| 6. $\text{NH}_4^{1+}$ $\text{P}^{3-}$   | 6. _____  |
| 7. $\text{Cu}^{1+}$ $\text{NO}_3^{1-}$  | 7. _____  |
| 8. $\text{Cu}^{2+}$ $\text{NO}_3^{1-}$  | 8. _____  |
| 9. $\text{Pb}^{4+}$ $\text{O}^{2-}$     | 9. _____  |
| 10. $\text{Li}^{1+}$ $\text{CO}_3^{2-}$ | 10. _____ |

For each of the following compounds, write... A) the symbols of the ions in the compound, and B) the number of each ion in one molecule of that compound.

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|--|-----------|
| 11. $\text{CaI}_2$                     | 11. _____ |
| 12. $\text{Na}_2\text{CO}_3$           | 12. _____ |
| 13. $\text{Ga}(\text{ClO}_3)_3$        | 13. _____ |
| 14. $\text{CuF}_2$                     | 14. _____ |
| 15. $(\text{NH}_4)_3\text{PO}_4$       | 15. _____ |
| 16. $\text{FeSO}_4$                    | 16. _____ |
| 17. $\text{Mg}(\text{NO}_3)_2$         | 17. _____ |
| 18. $\text{NH}_4\text{NO}_2$           | 18. _____ |
| 19. $\text{KC}_2\text{H}_3\text{O}_2$  | 19. _____ |
| 20. $\text{Na}_2\text{Cr}_2\text{O}_7$ | 20. _____ |