

HOMEWORK 5-10**VOCABULARY****Complete each sentence.**

1. The measure of the ability of an atom in a chemical compound to attract electrons is called

_____.

2. In general, this ability to attract electrons _____ across each period.

3. Many noble gases cannot be assigned these values because _____

_____.

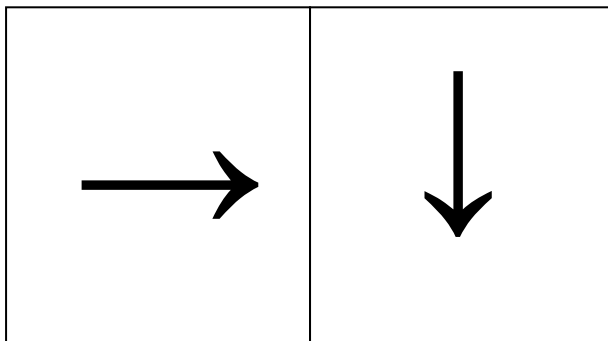
4. To determine the electron configuration of an ion, you should consider the reverse of

_____.

5. Valence electrons are electrons in an atom that _____

GRAPHIC ORGANIZER

These two boxes are representations of the periodic table of elements. Look at the way the arrows are pointing in the boxes. On a separate sheet of paper, list the trends that are represented by each of the arrows, such as *increasing atomic radius*, *decreasing electronegativity*, and so on.

**STANDARDIZED TEST PREP****Circle the letter of the best answer.**

- Electronegativity values are based on
 - quantities relative to the electronegativity of fluorine.
 - the atomic masses of the elements.
 - the ionization energy and electron affinities of the elements.
 - locations in the periodic table.
- Which of the following statements is true?
 - The atomic radii of the *d*-block elements generally increase across the periods.
 - Electrons in the lowest occupied sublevels are always removed first.
 - Periodic properties of the *d*-block elements vary with less regularity than main-group elements.
 - Noble gases have relatively low electronegativities compared to other elements.