

name: \_\_\_\_\_

date: \_\_\_\_\_

## writing formulas for ionic compounds - practice problems - B

**Directions:** Complete the following table. The first one has been completed as an example. Use a periodic table and polyatomic ion chart as needed.

cation	anion	ionic formula
cesium	thiocyanate	<b>CsSCN</b>
gold (III)	hydrogen phosphate	
nickel (I)	nitrate	
rubidium	nitrite	
tin (IV)	carbonate	
beryllium	borate	
ammonium	arsenite	
silver	sulfate	
copper (II)	sulfite	
mercury (II)	thiosulfate	
aluminum	chlorate	
iron (II)	molybdate	
manganese (III)	superoxide	
cobalt (I)	permanganate	
titanium	cyanide	
zinc	disulfide	
strontium	hydrogen carbonate	
potassium	chlorite	
lead (IV)	acetate	
lithium	hydrogen phosphite	

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cesium	thiocyanate	<b>CsSCN</b>
gold (III)	hydrogen phosphate	Au <sub>2</sub> (HPO <sub>4</sub> ) <sub>3</sub>
nickel (I)	nitrate	NiNO <sub>3</sub>
rubidium	nitrite	RbNO <sub>2</sub>
tin (IV)	carbonate	Sn(CO <sub>3</sub> ) <sub>2</sub>
beryllium	borate	Be(BO <sub>3</sub> ) <sub>2</sub>
ammonium	arsenite	(NH <sub>4</sub> ) <sub>3</sub> AsO <sub>3</sub>
silver	sulfate	Ag <sub>2</sub> SO <sub>4</sub>
copper (II)	sulfite	CuSO <sub>3</sub>
mercury (II)	thiosulfate	HgS <sub>2</sub> O <sub>3</sub>
aluminum	chlorate	Al(ClO <sub>3</sub> ) <sub>3</sub>
iron (II)	molybdate	FeMoO <sub>4</sub>
manganese (III)	superoxide	Mn <sub>2</sub> (O <sub>2</sub> ) <sub>3</sub>
cobalt (I)	permanganate	CoMnO <sub>4</sub>
titanium	cyanide	Ti(CN) <sub>2</sub>
zinc	disulfide	ZnS <sub>2</sub>
strontium	hydrogen carbonate	Sr(HCO <sub>3</sub> ) <sub>2</sub>
potassium	chlorite	KClO <sub>2</sub>
lead (IV)	acetate	Pb(C <sub>2</sub> H <sub>3</sub> O <sub>2</sub> ) <sub>4</sub>
lithium	hydrogen phosphite	Li <sub>2</sub> HPO <sub>3</sub>