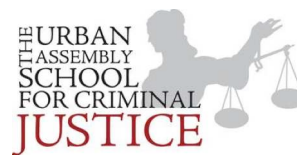


Name: _____ Date: _____

Chemistry ~ Ms. Hart

Class:

Anions or Cations



6.5 Balancing Equations

Directions: balance each of the equations below. Use a sheet of loose leaf if you need more space.

1. _____ Sb + _____ Cl ₂ → _____ SbCl ₃	18. _____ NaBr + _____ Cl ₂ → _____ NaCl + _____ Br ₂
2. _____ Mg + _____ O ₂ → _____ MgO	19. _____ Zn + _____ AgNO ₃ → _____ Zn(NO ₃) ₂ + _____ Ag
3. _____ CaCl ₂ → _____ Ca + _____ Cl ₂	20. _____ Sn + _____ Cl ₂ → _____ SnCl ₄
4. _____ NaClO ₃ → _____ NaCl + _____ O ₂	21. _____ Ba(OH) ₂ → _____ BaO + _____ H ₂ O
5. _____ Fe + _____ HCl → _____ FeCl ₂ + _____ H ₂	22. _____ Mg(OH) ₂ + _____ HCl → _____ MgCl ₂ + _____ H ₂ O
6. _____ CuO + _____ H ₂ → _____ Cu + _____ H ₂ O	23. _____ Na ₂ CO ₃ + _____ HCl → _____ NaCl + _____ H ₂ CO ₃
7. _____ Al + _____ H ₂ SO ₄ → _____ Al ₂ (SO ₄) ₃ + _____ H ₂	24. _____ NH ₄ NO ₂ → _____ N ₂ + _____ H ₂ O
8. _____ MgBr ₂ + _____ Cl ₂ → _____ MgCl ₂ + _____ Br ₂	25. _____ N ₂ + _____ O ₂ → _____ N ₂ O ₅

9. ____ SnO_2 + ____ C $\rightarrow$ ____ Sn + ____ CO	26. ____ MgCO_3 $\rightarrow$ ____ MgO + ____ CO_2
10. ____ $\text{Pb}(\text{NO}_3)_2$ + ____ H_2S $\rightarrow$ ____ PbS + ____ HNO_3	27. ____ KBr + ____ Cl_2 $\rightarrow$ ____ KCl + ____ Br_2
11. ____ HgO $\rightarrow$ ____ Hg + ____ O_2	28. ____ Zn + ____ CuSO_4 $\rightarrow$ ____ Cu + ____ ZnSO_4
12. ____ KClO_3 $\rightarrow$ ____ KCl + ____ O_2	29. ____ P + ____ O_2 $\rightarrow$ ____ P_4O_6
13. ____ N_2 + ____ H_2 $\rightarrow$ ____ NH_3	30. ____ SrBr_2 + ____ $(\text{NH}_4)_2\text{CO}_3$ $\rightarrow$ ____ SrCO_3 + ____ NH_4Br
14. ____ AgNO_3 + ____ $(\text{NH}_4)_2\text{CrO}_4$ $\rightarrow$ ____ Ag_2CrO_4 + ____ NH_4NO_3	31. ____ Li + ____ O_2 $\rightarrow$ ____ Li_2O
15. ____ K + ____ H_2O $\rightarrow$ ____ KOH + ____ H_2	32. ____ ZnCl_2 + ____ KOH $\rightarrow$ ____ $\text{Zn}(\text{OH})_2$ + ____ KCl
16. ____ Al + ____ $\text{Pb}(\text{NO}_3)_2$ $\rightarrow$ ____ $\text{Al}(\text{NO}_3)_3$ + ____ Pb	33. ____ Fe_3O_4 + ____ H_2 $\rightarrow$ ____ Fe + ____ H_2O
17. ____ Fe + ____ O_2 $\rightarrow$ ____ Fe_3O_4	34. ____ $\text{Pb}(\text{NO}_3)_2$ $\rightarrow$ ____ Pb + ____ NO_2 + ____ O_2