

Ionic VS Molecular Compounds Practice Worksheet

REMEMBER - Ion charge is only for ionic compounds; molecular do not form ions*

Example: CO

What elements are in this compound?

Are they metal or non-metal?

Is the compound ionic or molecular?

How many carbon atoms? _____

How many oxygen atoms? _____

Total number of atoms? _____

Name it:

1. magnesium nitride Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound formula: _____

Total # of atoms: _____

2. silicon dioxide Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound formula: _____

Total # of atoms: _____

3. CF_4 Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound name: _____

Total # of atoms: _____

4. S_4N_2 Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound name: _____

Total # of atoms: _____

5. Ga_2O_3 Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound name: _____

Total # of atoms: _____

6. dinitrogen trioxide Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound formula: _____

Total # of atoms: _____

7. beryllium nitride Ionic or molecular: _____

Ion charge (for all ions in compound): _____

Compound formula: _____

Total # of atoms: _____

8. CaCO_3 Total # of atoms: _____

9. $\text{HC}_2\text{H}_3\text{O}_2$ Total # of atoms: _____

10. Ion charge of Cu in Cu_3P _____

11. Ion charge of Mn in MnO_2 _____

12. Ion charge of Cr in CrCl_3 _____