

Chemistry A Ch. 5&6 Practice Test**Matching**

Match each item with the correct statement below.

- | | |
|----------------------|-----------------|
| a. electronegativity | f. periodic law |
| b. ionization energy | g. cation |
| c. atomic radius | h. period |
| d. metal | i. group |
| e. transition metal | j. electrons |

- ___ 1. horizontal row in the periodic table
- ___ 2. vertical column in the periodic table
- ___ 3. A repetition of properties occurs when elements are arranged in order of increasing atomic number.
- ___ 4. type of element that is a good conductor of heat and electric current
- ___ 5. type of element characterized by the presence of electrons in the *d* orbital
- ___ 6. one-half the distance between the nuclei of two atoms when the atoms are joined
- ___ 7. type of ion formed by Group 2A elements
- ___ 8. subatomic particles that are transferred to form positive and negative ions
- ___ 9. ability of an atom to attract electrons when the atom is in a compound
- ___ 10. energy required to remove an electron from an atom

Short Answer

- 11. Which group of elements in the periodic table is known as the alkali metals?

- 12. Which group in the periodic table is known as the noble gases?

- 13. An element has an atomic number of 80. How many protons and electrons are in an atom of the element?

Name: _____

ID: A

22. Elements in which the *d*-sublevel is being filled have the properties of

23. The alkali metals belong to the ____-block in the periodic table.

24. How do scientists determine the radius of an atom?

25. Write the chemical symbol and tell what kind of ion the following atoms will form:
B, I, Po, N, Be, F, Rb,

Numeric Response

26. How many electrons are in a rubidium ion (Rb^+)?
27. What is the usual charge on an ion from Group 7A?
28. How many electrons does the ion Ca^{2+} contain?

Essay

29. Describe the trends in the atomic size of elements within groups and across periods in the periodic table. Provide examples.

